

Endothermic or Exothermic Reaction Lab

Problem: Will a combination of Sodium Bicarbonate(NaHCO_3) and dilute Acetic Acid($\text{HC}_2\text{H}_3\text{O}_2$) produce an endothermic or exothermic chemical reaction (or not react)?

Hypothesis: (Form your own that speaks to the problem.)

Materials:

Erlenmeyer flask

250 ml beaker

Graduated cylinder

Triple-beam balance

Styrofoam cup

80 ml Acetic Acid

10 grams Sodium Bicarbonate

Timing Device (stopwatch or timer on cell phone)

Procedure:

Pour 80 ml of acetic acid in a 250 ml beaker.

Record the temperature in the beaker.

Add the Sodium Bicarbonate.

Record the temperature in the beaker at 5 Second intervals for 2 minutes.

Data: Present your data in chart and graph form.

Conclusion: Speak to your hypothesis....